pHooey! HW

Read and outline (Cornel Style) **Section 19.3** in your chemistry textbook. ****Your section outlines must be at least ½ page with 3-4 topic questions**** Then answer the following assessment questions:

1. What is the relationship between the pH of a solution and the concentration of hydrogen ions in the solution?
2. If you know the pOH of a solution, how can you determine its pH?
3. Use your calculator to figure out the pH of each of the following solutions:
    a. $[H^+] = 0.0014M$
    b. $[H^+] = 6.0 \times 10^{-8}M$
    c. $[H^+] = 4.2 \times 10^{-11}M$
    d. $[H^+] = 1.5 \times 10^{-1}M$
    e. $[H^+] = 2.4 \times 10^{-13}M$
    f. $[OH^-] = 5.0 \times 10^{-4}M$
4. Which of the following solutions in question 3 are acids?